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Describe the nuclear model of the atom, including the general location of the protons, neutrons, and electrons, the relative size of the nucleus compared to the size of the atom, and the modern description of the electron.

## *Chapter 18 Nuclear Chemistry*

Nuclear Chemistry. Radioactive decay. Fission. Chain reaction. Fusion. the emission of radiation from an unstable nucleus. a nuclear reaction where the nucleus of a large, unstable atom.... a reaction in which the products from one reaction are used to.... a nuclear reaction where two small nuclei combine, or fuse tog....

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Nuclear chemistry is the part of chemistry that studies the reactions that affect nucleons (particles in the nucleus). As such, radioactive decay (the spontaneous breakdown of the nucleus), nuclear...

*What is nuclear chemistry? | Study.com*

Proton = 1. Neutron = 1. Electron = 0.

Define nuclear binding energy and describe its trend. The very small amount of energy that is released when a nucleus is formed from nucleons. Its Trend is: as the nucleus increases in size and number of nucleons , the nuclear binding energy increases until it gets up to Iron (Fe).

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PLAY. Nuclear chemical reactions.  
reactions that occur because of changes in  
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chemical reactions occur?-atoms strive to  
become stable with an equal proton to  
neutron ratio

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1) Distinguish between nuclear fission and  
fusion. Provide AT LEAST 2 facts about

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each to support your response. ~ 4 pts 2)  
Name 3 ways that isotopes of an element differ. ~ 6 pts  
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25. Name 25 Class Section 25.2

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## *Unit 17 Nuclear Chemistry Study Guide Answers*

A combination of radiochemistry and radiation chemistry is used to study nuclear reactions such as fission and fusion. Some early evidence for nuclear fission was the formation of a short-lived radioisotope of barium which was isolated from neutron irradiated uranium ( $^{139}\text{Ba}$ , with a half-life of 83 minutes and  $^{140}\text{Ba}$ , with a half-life of 12.8 days, are major fission products of uranium).

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